

B. PHARM.
FIRST SEMESTER
PHARMACEUTICAL INORGANIC CHEMISTRY
BP104T [SPECIAL REPEAT]
[USE OMR SHEET FOR OBJECTIVE PART]

SET
A

Duration : 3 hrs.

Full Marks : 75

(PART-A: Objective)

Time : 30 min.

Marks : 20

Choose the correct answer from the following:

1×20=20

- According to Arrhenius concept an base is a substance which
 - Gives H⁺ ion
 - Gives OH⁻ ion
 - Hydronium ion
 - None
- Which compound produce if Al(OH)₃ gel reacts with gastric HCl?
 - Al₃O₃
 - AlCl₃
 - Al₃Cl₃
 - Al₃Cl
-is written detail study about drug which are mentioned in Pharmacopoeia.
 - Maps
 - Astronomical books
 - Monograph
 - Assay
- Inorganic compounds generally do not Contain..... atoms
 - N
 - C
 - O
 - S
- In limit test for iron , ferrous thioglycolate has stable pink to reddish purple color in which medium?
 - Acidic.
 - alkaline
 - Neutral.
 - None
- Washing soda is
 - Sodium carbonate
 - Sodium bicarbonate
 - Sodium sulphite
 - Sodalimne
- What is added in preparation of barium Sulphate reagent to prevent super Saturation
 - Ethanol
 - Barium Chloride
 - Potassium Sulphate
 - None
- Synonym of Magnesium Sulfate is
 - Baking soda
 - Fitkari
 - Epsom Salt
 - Milk of magnesia
- The basic objective of replacement therapy is
 - To restore the volume and composition of body fluids
 - To restore the volume only
 - To restore the composition of body fluids only
 - None of these

(PART-B :Descriptive)

Time : 2 hrs. 30 min.

Marks : 35

[Answer any seven (7) questions]

1. Briefly discuss about ORS with composition recommended by WHO and UNICEF for controlling diarrhoea. 1+1+2+
1=5
2. Define limit test. Explain the Principle and reactions involved in iron limit test I.P elaborating on the specific uses of each reagent used. 1+2+2
=5
3. Define P^H and derive the equation for P^H scale. 1+4=5
4. What is the chemical name and formula of 'green vitriol'? Indicate its used and explain principle of its assay. 2+3=5
5. Define isotopes. What are the properties of radioisotopes? Give two example and indicate their uses. 2.5+2.5
=5
6. Classify dental products. Define each class with a suitable example. 1+1+3
=5
7. Define and classify antidote. Give two examples. Explain treatment of cyanide poisoning. 1+1+1+
2=5
8. Define Emetics. Classify them with example. Write the MOA of Emetics. Write the molecular formula, synonym, preparation and uses of copper sulphate. 1+1+1+
2=5
9.
 - a. Define Half Life of compound. 1+1+1+
 - b. Write the molecular formula and synonym of Calcium Hydroxide. 1+1=5
 - c. Define astringent and haematinics with example.
 - d. Write the role of lead acetate cotton wool in the limit test for arsenic.
 - e. Write the Molecular formula and uses of Potash Alum.

(PART-C: Long type questions)

[Answer any two (2) questions]

1. Define Acid and Base according to Traditional, Arrhenius and Bronsted Lowry Concept. Derive Henderson Hassalbalch equation for weak acid. 6+4=10

2. Define Acidifiers. Classify them with example. Write the Monograph of Ammonium Chloride. 1+2+7
=10

3. Explain the function of major physiological ions. Write a note on electrolyte replacement therapy. Explain with suitable equation, the principle involved in the assay of NaCl I.P 3+3+4
=10

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